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so many fake sites. this is the first one which worked! Many thanks

Lab: Single and Double Replacement Reactions

Purpose: To predict products and observe various single and double replacement reactions

Hypothesis: Single Replacement Reactions: Use the **activity series** (a list – found and your understanding of single replacement reactions to predict products for the following reactions. If no reaction will occur, then write "no rxn" on the product side of the equation.

1) $HCl + Ag \rightarrow$ _____

2) $HCl + Cu \rightarrow$ _____ (turn copper coated bar of wire into compound)

3) $HCl + Fe \rightarrow$ _____ (turn iron bar to iron (II) ion compound)

4) $HCl + Zn \rightarrow$ _____

Hypothesis: Double Replacement Reactions: Use the **solubility chart** and your understanding of double replacement reactions to predict products for the following reactions. If both of the possible products are aqueous, then write "no rxn" on the product side of the equation. Otherwise, clearly state which product will be aqueous and which will form a solid precipitate.

1) $Na_2CO_3(aq) + Ca(NO_3)_2(aq) \rightarrow$ _____

2) $NaOH(aq) + Fe(NO_3)_3(aq) \rightarrow$ _____

3) $Na_3PO_4(aq) + AgNO_3(aq) \rightarrow$ _____

4) $Ag_2SO_4(aq) + Zn(NO_3)_2(aq) \rightarrow$ _____

Procedures:

1. For each single replacement reaction, place a sample of each metal in your well plate and then place 5-6 drops of HCl (aq) on top of it. Record your observations.
2. For each double replacement reaction, place two drops of each solution carefully on the transparency, covering each "X" appropriately over your "Super High-Tech Patented Reaction Surface." Record your observations.
3. If a precipitate formed, clearly write the formula for the precipitate on the line provided.
4. Clean up your workstation according to your instructors directions and answer all questions.

Observations/Data: Table 1: Single Replacement Reaction Observations

	HCl (aq)
Ag	1 observation:
Cu	2 observation:
Fe	3 observation:
Zn	4 observation:

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