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so many fake sites. this is the first one which worked! Many thanks

COMMON CHEMICAL REACTION - THIS CAN BECOME YOUR EXPECTED YIELD

Actual yield - In reality, reactions don't work perfectly and some product either doesn't get made or gets lost along the way (e.g. left in a filter, products turned back into reactants, some reactants reacted, etc.). Actual yield is the actual mass that you measure during lab of product formed. It should be LESS than the theoretical yield.

Question - Why should the actual yield be less than the theoretical, not more?

Processes that commonly decrease actual yield

- Distillation** - liquids boiled to separate them from other chemicals
- Recrystallization** - chemicals are dissolved and then allowed to turn back into crystals
- Side reactions** - Other reactions that can occur with the reactants used that will make other products, instead of what you wanted.

Percent Yield is the ratio relating the actual yield of a reaction to its theoretical yield to describe the efficiency of the reaction.

$$\text{Percentage Yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100\%$$

Example 1 - You calculate that you should get 5.4 grams of copper in your reaction, however you only measure 2.3 grams of copper at the end.

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