

Download File PDF Specific Heat Worksheets And Answers

#Jenny



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#Markus Jensen



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so many fake sites. this is the first one which worked! Many thanks

• WS 1.4 Specific Heat & Calorimetry
(Show all work neatly) $Q = m \cdot c \cdot \Delta T$

Specific Heats	
Substance	c (J/g°C)
water	4.184
ethanol	2.42
graphite	0.71
diamond	0.50
iron	0.449
copper	0.385
silver	0.235
gold	0.129
tin	0.213

- How much heat is required to raise the temp of 854 g of water from 34.5°C to 89.7°C? Ans: _____
- How much heat is required to raise the temp of 504 g of glass from 34.5°C to 89.7°C? Ans: _____
- If 7200 J were added to 152 g of ethanol, its temp would go up by how much? Ans: _____
- 16.25 g of water at 54.0°C releases 402 J. What will be its final temp? (See the example on the back page of this book for help.) Ans: _____
- 897 J are added to a 36.8 g of benzene and the temp increases from 22.5°C to 34.7°C. Determine benzene's specific heat. Ans: _____
- 25 copper pennies (each weighing 3.12 g) are placed in 36.0 g of ethanol at room temp (22.1°C). How much heat will it take to raise the temperature up to 65.4°C? (See the example on the back page of this book for help.) Ans: _____
- What mass of 50.0°C water must be added to 500 g of 21.0°C water to make the final temp of both come out to be 29.9°C? Ans: _____
- What mass of 54.0°C gold must be added to 400 g of 21.0°C water to make the final temp of both come out to be 26.9°C? Ans: _____
- A 525 g brass rod at 170.0°C is placed in a cup containing 142 g of 24.3°C water. The final temp comes out to be 37.4°C. Determine brass's specific heat. Ans: _____
- 100.0 g of water at 20.0°C are mixed with 200.0 g of copper at 60.0°C. What will the final temp come out to be? Ans: _____

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